

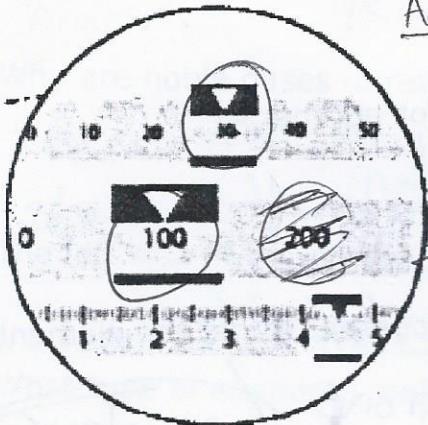
1st Six Weeks Test Study Guide

1. List the function of the following science equipment.

a. Beaker- measure large amounts of liquid (volume),
 b. Graduated Cylinder- measure small amount of liquid
 c. Hot Plate- heat up liquid in lab
 d. Lab Journal/ notebook- Record ~~data~~ and analyze data
 e. Meter Stick- measure length or distance cm/m.
 f. Safety goggles- to protect eyes from harmful chemicals
 g. Thermometer- measure temperature $^{\circ}\text{C/F}$

2. Calculate the mass in grams for the following examples:

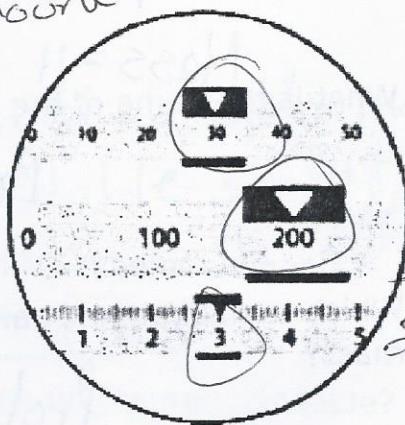
Add larger amount



a.

$$\text{Mass} = \underline{134.5} \text{ g}$$

100
30
4.5
134.5

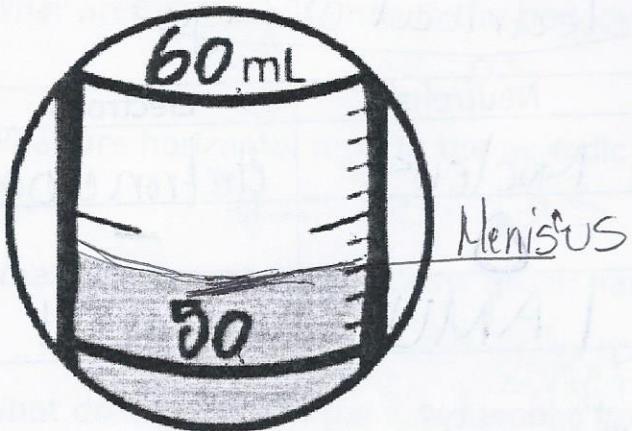


b.

$$\text{Mass} = \underline{233} \text{ g}$$

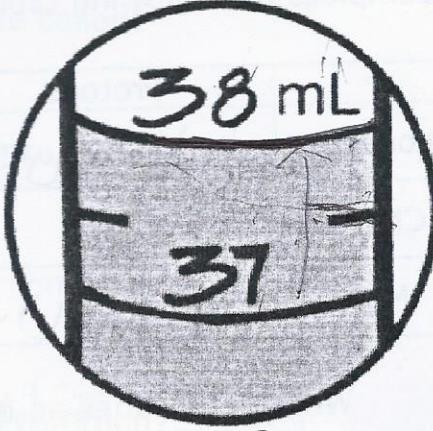
200
30
3
233

3. What is the volume in mL for the following examples:



a.

$$\text{Volume} = \underline{53} \text{ mL}$$

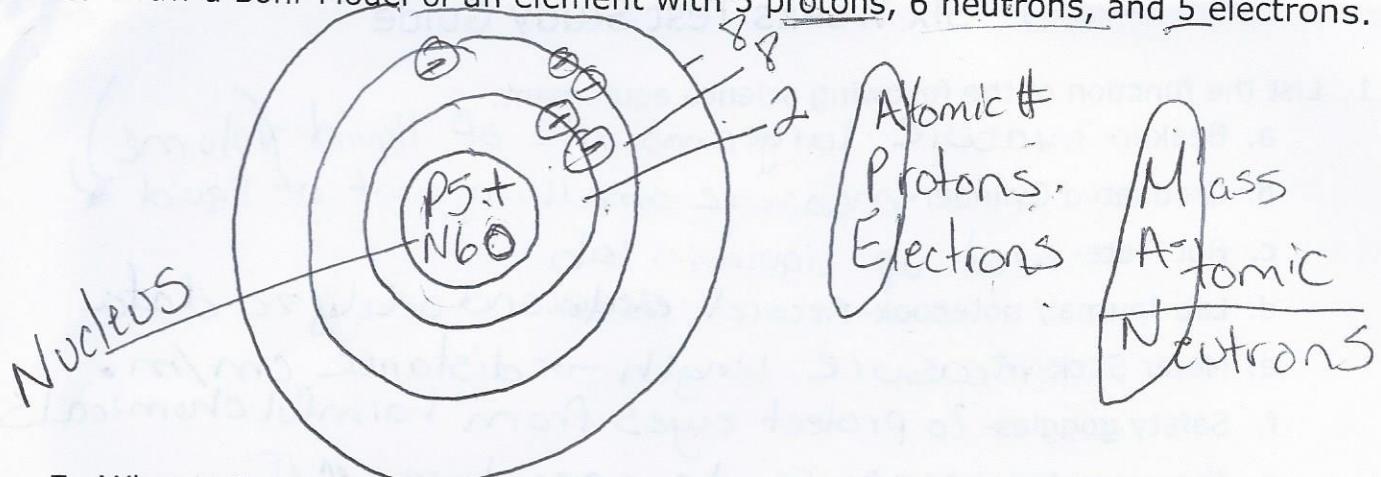


b.

$$\text{Volume} = \underline{38} \text{ mL}$$

Name _____ Date _____

4. Draw a Bohr Model of an element with 5 protons, 6 neutrons, and 5 electrons.



5. What is the atomic number and atomic mass of the element you drew for number 4?

$$A=5$$

$$\text{Mass}=11$$

6. What is the name of the element you drew for number 4?

Boron

7. Which two subatomic particles do you ADD together to get the element's atomic mass?

Proton & Neutrons ~~Nucleus~~

8. In a neutral atom, which two subatomic particles are always equal?

Protons + & Electrons ~~E~~ opposite attract

9. Complete the following table.

electrical charge

	Protons	Neutrons	Electrons
Location	Nucleus	Nucleus	electron cloud
Charge	+	0	-
Mass	1 AMU	1 AMU	0 AMU

11. What determines an element's identity?

Protons *

Atomic #

12. The atomic number of an element is equal to which subatomic particle?

Atomic
Protons
Electrons

Name _____

Date _____



13. What are valence electrons?

-The electrons found in the outermost shell

-Find them by looking group #

14. What determines an element's chemical property? How?

By the groups according to valence electrons determine reactivity

15. List the 6 noble/inert gases.

Helium	Krypton
Neon	Xenon
Argon	Radon

16. Why are noble gases unreactive?

Because they are full on their outermost shell so therefore they are stable & ~~want to be left alone~~

17. The left side of the periodic table consists of which type of elements?

Metals



18. What type of elements make up the zig-zag line on the periodic table?

metalloids

19. The right side of the periodic table consists of which type of elements?

Non-Metals

20. What are vertical columns in the periodic table called?

↑ groups

21. What are horizontal rows in the periodic table called?

→ Periods

22. What do elements in the same group have in common?

* Valence electrons
Share properties, reactivity

23. What do elements in the same period have in common?

same energy levels

